



## COURSE DESCRIPTION CARD - SYLLABUS

Course name

Kinetyka chemiczna i elektrochemia II (Chemical kinetics and electrochemistry II)

### Course

Field of study

Year/Semester

Technologia chemiczna (Chemical Technology)

II/4

Area of study (specialization)

Profile of study

-

general academic

Level of study

Course offered in

First-cycle studies

Polish

Form of study

Requirements

full-time

elective

### Number

#### of hours

Lecture

Laboratory classes

Other (e.g. online)

0

15

0

Tutorials

Projects/seminars

0

0

#### Number of credit points

2

### Lecturers

Responsible for the course/lecturer:

Responsible for the course/lecturer:

Prof. Andrzej Lewandowski

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Wydział Technologii Chemicznej

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### Prerequisites

Students:

have knowledge in general chemistry (writing chemical reactions, converting concentrations, knowledge of laboratory glassware and basic laboratory equipment).

have knowledge in mathematics and physics enabling the introduction of problems in physical chemistry (basic laws of physics, differential calculus).

are able to prepare solutions of specific concentrations.



are aware of further development of their competences.

### Course objective

To familiarise students with the practical application of advanced problems in physical chemistry and electrochemistry at the academic level in the field of: chemical kinetics (simple and complex reactions in simple laboratory experiments), catalysis, corrosion and the theory of strong and weak electrolytes.

### Course-related learning outcomes

#### Knowledge

Students will be able to characterise, list and identify simple and complex reactions, define homo- and heterogeneous catalysis and conductivity. K\_W03, K\_W10

Students will be able to define and explain selected problems in electrochemistry: mechanism and types of corrosion. K\_W03, K\_W10

Students will be able to define and explain problems in kinetics of complex reactions (oscillatory and catalytic reactions). K\_W03, K\_W10

#### Skills

Students will be able to obtain information from literature, databases and other sources; interpret it as well as draw conclusions and formulate and substantiate opinions. K\_U01

Students will be able to work individually and as part of a team; estimate the time needed to complete the assigned task. K\_U2

Students will be able to apply the principles of thermodynamics in the implementation of chemical processes. K\_U23

Students will have the self-study skills in the subject. K\_U05

Students will be able to elaborate, describe and present results of an experiment or theoretical calculations. K\_U09

Students will be able to distinguish between types of chemical reactions and to select them for specific chemical processes. K\_U18

#### Social competences

Students will understand the need for further training and developing their professional competences. K\_K01

Students will be able to properly prioritise the task. K\_K04

### Methods for verifying learning outcomes and assessment criteria

Learning outcomes presented above are verified as follows:



Laboratory classes: The course passing is based on points received for the individual exercise description. Passing exercises from 56% .

### Programme content

Laboratory classes:

Rate of chemical reaction. Rate constant. Order of the Chemical Reaction. Molecularity of the Chemical Reaction. Zero, first, second, and third-order rate equations.

Complex Reactions: reversible, parallel, competitive, consecutive. Inductive. Oscillatory Reactions.

Catalysis. Heterogeneous and Homogenous Catalysis. Mechanism of catalyst operation. Types of homogeneous catalysts in liquid solution. The dependence of rate of catalysis on the amount of catalyst. Acid-base catalysis. The principle of operation of heterogeneous solid catalysts. Catalyst supports (powder and monolithic).

Chemical and electrochemical corrosion (examples). Mechanism of electrochemical corrosion processes. Protecting from Corrosion. Chemical and electrochemical depositions of metal coating. Standard Electrochemical potentials series. Electrolysis

Theory of the strong and weak electrolytes. Conductometry. Conductivity. Specific conductivity. Methods of conductivity measurements. Law of independent migration of ions – Kohlrausch's law. Limiting molar conductivity determination for weak and strong electrolytes. Conductivity in non-aqueous electrolytes

### Teaching methods

Laboratory classes- practical method - laboratory exercises. Planning, execution and analysis of the results of physicochemical experiment.

### Bibliography

Basic

1. K. Pigoń, Z. Ruziewicz, Chemia Fizyczna, PWN Warszawa 2013
2. P. Atkins, Chemia Fizyczna, PWN Warszawa 2019
3. A. Molski, Wprowadzenie do kinetyki chemicznej, WNT Warszawa 2000
4. L. Sobczyk, Eksperymentalna Chemia Fizyczna, PWN Warszawa 1982

Additional

1. P. Atkins, Podstawy Chemii Fizycznej, PWN Warszawa 1999
2. L. Sobczyk, A. Kiszka, Chemia fizyczna dla przyrodników, PWN Warszawa 1977
3. H. Buchnowski, W. Ufnalski Wykłady z chemii fizycznej, WNT Warszawa 1998



4. Instrukcje do ćwiczeń laboratoryjnych z chemii fizycznej

**Breakdown of average student's workload**

	Hours	ECTS
Total workload	50	2,0
Classes requiring direct contact with the teacher	25	1,0
Student's own work ( preparation for laboratory classes, preparation of the report. ) <sup>1</sup>	25	1,0

<sup>1</sup> delete or add other activities as appropriate